

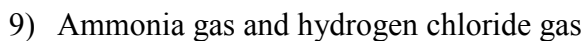
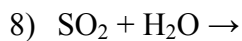
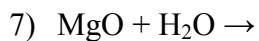
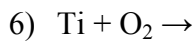
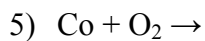
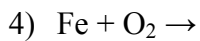
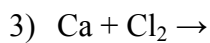
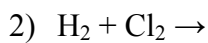
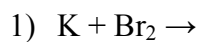
Name \_\_\_\_\_

AP Chemistry

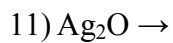
## Reaction Prediction

Write the balanced molecular equation along with the net ionic equation.

### *Synthesis Reactions*



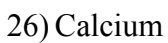
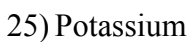
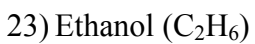
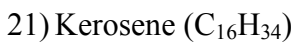
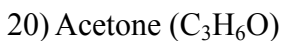
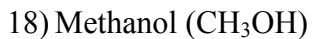
### *Decomposition*



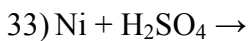
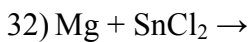
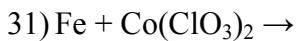
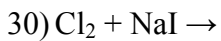
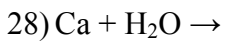


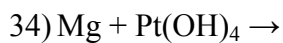
17) Mercury (II) oxide is a bright red powder. It decomposes on heating.

*Combustion*

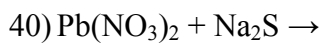
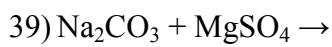
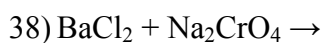
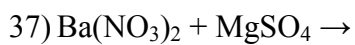
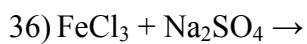
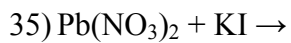


*Oxidation-reduction*

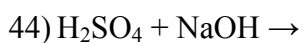
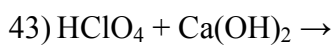
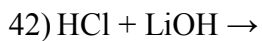




*Double Displacement (Precipitation)*



*Double Displacement (Acid-Base)*



45) Solid cadmium sulfide reacts with an aqueous solution of sulfuric acid

46) Solid magnesium carbonate reacts with an aqueous solution of perchloric acid

47) Solid  $\text{CaCO}_3$  reacts with an aqueous solution of nitric acid

48) Solid iron (II) sulfide reacts with an aqueous solution of hydrobromic acid